

O & A - Levels: Chemistry | Maths

	O & A - Levels. Chemistry   Mauis	
	Name: Back Titration - Suggested Solution Date:	
***************************************	Reaction Equations:	
	NH4 Clags + Na OH cags > Na Clags + H2 O(1) + NH3 cg 2Na OH cags + H2 SO4 cags -> Na 2 SO4 cags + 2H2 O(2)	)
	2 Na O H cap + H2 SO4 cap -> Na2 SO4 cap + 2 H2 O(2)	
	Key Concepts:	
	known excess amount of NaOH added initially	
<b>O</b>		
	amount of NaOH amount of NaOH	
	that has reacted that has reacted	
	with NH4Cl with H2SO4	
	in first reaction in second reaction	
	i·e· unreacted excess amount	
	of NaOH after first reaction.	
	Calculations:	
	Amount of H2SO4 = 50 × 0.250 = 0.0125 mol	
0	HMOUNC 07 H2304 - 1000 101200 - 00(23110)	
	Amount of NaOH = Amount of $H_2SO_4 \times 2$ (unreacted excess) = $0.0125 \times 2 = 0.025$ mol	
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	Amount of NaOH added initially = 100 x 1:00	
	Amount of NaOH added initially = 1000 x 1.00 = 0.1 mol	
	Amount of NaOH that has reacted with NH4Cl	
	= Amount of NaOH added initially - Amount of unreacted excess	کڌ
	= Amount of NaOH added initially - Amount of unreacted exces = 0.1 - 0.025 = 0.075 mol NaOH	
	<i>!</i>	
	:. Amount of NH4Cl = 0.075 mol > Mass of NH4Cl = 0.075 x 53.5	5
	:. Amount of NH4Cl = 0.075 mol > Mass of NH4Cl = 0.075 x 53.5  = 4.019(35.4)	

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